

Ph Of Naoh Solution

Ph Of Naoh Solution - If you have dissolved 1 g of NaOH in enough water to make a total of 250 ml of solution, calculate the number of moles of solute present by dividing the mass of NaOH by the molecular mass of the compound. The molecular mass of NaOH is 40, so work out $1 \div 40 = 0.025$. Next, calculate the number of liters of solution present. Since NaOH is a strong base, .1 mol of it will dissociate into .1 mol Na⁺ and OH⁻. By definition $[H^+] \times [OH^-] = K_w$ where K_w , the ionization constant of water, is 10^{-14} . This is why the pH scale is 0–14 and also why looking at the exponent H⁺ concentration will generally tell you approximately the pH.